## Exercise 6.2 Reactions of Metals with Acids (Including Water)

This exercise assumes that you can do question 1 on Exercise 6.1.

1.	Describe what happens when a metal	l reacts	with H <sup>+</sup> .	Identify both	products.
2. (a)	What is left after one H <sup>+</sup> is removed <i>More properly, "What is the conjugo</i> HCl (b) HBr			_	
3. (a)	List all products formed in each of the Leave ions formed as free ions.  Na reacts with H <sup>+</sup> (aq)	e follow	J	tions. with H <sup>+</sup> (aq)	
(c)	Li reacts with HCl(aq)	(d)	Mg react	ts with HCl(ac	q)
(e)	K reacts with H <sub>2</sub> O	(f)	Ba reacts	s with H <sub>2</sub> O	
4. (a)	Write a balanced chemical equation in <i>Include states of matter</i> .  Na reacts with H <sup>+</sup> (aq)			llowing reactions with H <sup>+</sup> (aq)	ions.
(c)	Li reacts with HCl(aq)	(d)	Mg react	ts with HCl(ac	q)
(e)	K reacts with H <sub>2</sub> O	(f)	Ba reacts	s with H2O	