

**Exercise 6.5**  
**Stoichiometry Practice**

1. Suppose that you want to make 12 g of lithium oxide. What are the minimum masses of lithium and oxygen you will need?
2. Potassium metal (0.25 g) is dropped into a beaker of water (100. mL). What is the concentration of the resulting KOH(aq) solution?  
*Assume that the volume of liquid in the beaker does not change as a result of this reaction.*  
*The density of water is 1.00 g/mL.*
3. A piece of magnesium with a mass of 185 mg is dropped into a beaker containing 75 mL of 1.25 M HCl(aq). Once the reaction is complete, what is the concentration of HCl(aq) remaining in the beaker?